NEW TRUST SCIENCE ACADEMY

Chemistry	Unit 1
Name:	Marks

Chemistry 9th Unit # 1

Q1. Circle the correct answer. 1x9=9

i. The third abundant gas found in the oceans is:
(a)carbon monoxide (b) oxygen (c)hydrogen (d) clorine

ii. One amu (atomic mass unit) is equivalent to:
(a) $1.66x10^{-24}$mg (b) $1.66x10^{-23}$g (c) $1.66x10^{-24}$g (d) $1.66x10^{-24}$kg

iii. All of the followings are ri-atomic molecule except:
(a) H_2 (b) O_3 (c) H_2O (d) CO_2

iv. The mass of one molecule of water is:
(a) 18 amu (b) 18 g (c) 18 mg (d) 18 kg

v. The number of protons in magnesium?
(a) 10 (b) 12 (c) 9 (d) 8

vi. Molar mass is usually expressed in grams. Which one of the following is molar mass of O_2 in amu
(a)32 amu (b)$53.12x0^{-24}$ amu (c)$1.92\ x0^{-25}$ amu (d)$192.64\ x0^{-25}$ amu

vii. How many members of moles are equivalent to 8 grams of CO_2?
(a)1.05 (b)0.18 (c)0.21 (d)0.24

viii. Molar mass of nitric acid is:
(a) 63 g (b) 63 mg (c) 63 amu (d) 63 kg

ix. The valancy of of oxygen is:
(a) 3 (b) 2 (c) 1 (d) 0

2. Give short answers to following. 2x8=16

I. Give significance of chemical formulae.
II. Differentiate b/w diatomic and triatomic molecules with examples.
III. Define analytical and industrial chemistry.
IV. Differentiate b/w Atom and Ion.
V. Calculate number of moles in 6g of water.
VI. Differentiate b/w molecular mass and formula mass.
VII. What is the molecular mass of nitric acid.
VIII. Differentiate b/w homoatomic and heteroatomic moleclues with examples?

3. Explain the following.

a) differentiate b/w mixtures and compounds? (5)
b) Calculate the number of moles, number of molecules and number of atoms present in 6 grams of water.

TALHA

BEST OF LUCK...!