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| <b>Chemistry-9<sup>th</sup></b> | <b>Unit 3</b> |
| <b>Name:</b>                    | <b>Marks</b>  |

**chemistry-9<sup>th</sup>**

**Q1. Circle the correct answer. (1x8=8)**

- i. The amount of energy given out when an electron is added to an atom is called:  
(a) Lattice energy      (b) ionization energy  
(c) electronegativity   (d) electron affinity
- ii. Mendeleev Periodic Table was based upon the:  
(a) Electronic configuration      (b) atomic energy  
(c) atomic number      (d) mass number
- iii. Long form of Periodic Table is constructed on the basis of:  
(a) Mendeleev postulate      (b) atomic number  
(c) Atomic mass      (d) mass number
- iv. 4<sup>th</sup> and 5<sup>th</sup> period of the long form of Periodic Table are called:  
(a) Short periods      (b) normal periods  
(c) long periods      (d) very long periods
- v. The valence of group 2 is:  
(a) 4      (b) 1      (c) 3      (d) 2
- vi. How many elements are there in 4<sup>th</sup> period?  
(a) 02      (b) 08      (c) 18      (d) 32
- vii. The atomic no. of fluorine is  
(a) 3      (b) 4      (c) 5      (d) 9
- viii. Group 17<sup>th</sup> belongs to:  
(a) Halogen      (b) Noble gases  
(c) alkali metal      (d) none of these

**2. Give short answers to following. (2x6=12)**

- I. What are the defects of Mendeleev's periodic table?
- II. Why law of octaves was failed?
- III. Why atomic number is a more fundamental property than atomic size?
- IV. Why shielding effect increases along a group?
- V. Give trend of electron affinity along groups.
- VI. Why ionization energy increases from left to right in a period?

**3. Explain the following.**

- a) Atomic size and atomic radius with example and periodic trend. (5)

**BEST OF LUCK...!**